

**MARK SCHEME for the October/November 2011 question paper  
for the guidance of teachers**

**0620 CHEMISTRY**

**0620/33**

Paper 3 (Extended Theory), maximum raw mark 80

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

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- 1 (a) (i) Cs / Fr [1]
- (ii) Br [1]
- (iii) U / Pu / Th [1]
- (iv) I or At [1]
- (v) As [1]
- (vi) He / Ne / Ar / Kr / Xe [1]
- (b) (i) GeO<sub>2</sub> / GeO [1]
- (ii) TeBr<sub>2</sub> / TeBr<sub>4</sub> [1]
- (c) (i) Sr<sup>2+</sup> [1]
- (ii) F<sup>-</sup> [1]
- 2 (a) (i) molecule / unit / simple compound / building block **and** used to make a polymer / big molecule / long chain / macromolecule [1]
- formation of a polymer / big molecule / long chain / macromolecule **or** joining of monomers **and** elimination / removal / formation of a simple or small molecule / H<sub>2</sub>O / HCl [1]
- note:** two points needed for 1 mark in both parts
- (ii) -O- linkage [1]  
three correct monomer units [1]  
continuation [1]
- (b) (i) catalyst **and** from living organism [1]  
**accept:** biological catalyst / protein catalyst
- (ii) enzyme denatured / destroyed [1]
- (iii) chromatography [1]  
locating agent / description of locating agent [1]  
measure R<sub>f</sub> / compare with standards [1]

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- 3 (a) sodium hydroxide solution [1]  
warm [1]  
(only) ammonium phosphate gives off ammonia / gas (which will turn red litmus paper blue) [1]  
**or:**  
sodium hydroxide solution [1]  
dissolve fertiliser in water [1]  
 $\text{Ca}^{2+}$  gives (white) ppt [1]  
**or:**  
flame test [1]  
 $\text{Ca}^{2+}$  brick red / orange / orange-red [1]  
 $\text{NH}_4^+$  no colour [1]
- (b) iron catalyst [1]  
pressure 150–300 atmospheres [1]  
temperature 370–470 °C [1]  
 $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$  [1]  
**note:** units required for temperature and pressure
- (c) potassium / K [1]
- (d) (i) needs to be soluble / in solution (to be absorbed by plants) [1]  
(ii) base [1]  
proton acceptor [1]
- (e) plant growth depends on soil acidity or pH / plants have optimum pH (for growth) [1]  
add  $\text{Ca}(\text{OH})_2$  / CaO /  $\text{CaCO}_3$  / lime / slaked lime / quicklime / limestone [1]
- 4 (a) (i) alloy / mixture [1]  
iron and carbon / another metal or element etc. [1]  
(ii) electron loss [1]
- (b) electrons move from / lost from Mg [1]  
to steel / iron [1]
- (c) (i)  $2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$  [2]  
not balanced = 1

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- (ii) sacrificial protection – is a cell [1]  
cathodic protection – is electrolysis NOT electrical cell [1]  
**or:**  
sacrificial protection – electrons from more reactive metal [1]  
cathodic protection – electrons from battery etc. [1]  
**or:**  
sacrificial protection – does not need or use power / battery / electricity / electrical cell [1]  
cathodic protection – does [1]  
**or:**  
sacrificial protection uses up / needs a sacrificial / more reactive metal [1]  
cathodic protection doesn't [1]
- 5 (a) light / UV / sun / sunlight / solar energy [1]  
starts / initiates / speeds up [1]
- (b) (i) 0.03% – 1(%) carbon dioxide [1]  
**accept:** less than 1(%)  
20% – 21(%) oxygen [1]
- (ii) remove carbon dioxide from atmosphere [1]  
produce oxygen [1]  
any **two** from:  
photosynthesis  
chlorophyll / chloroplast  
light / sun / sunlight / UV / photochemical  
formed carbohydrates / glucose / sugar(s) [2]
- (c) reaction is photochemical / needs light [1]  
(light) causes formation of silver / silver ions reduced [1]  
(on formation of silver) goes black [1]  
no light still silver(I) bromide / stays white / no reaction [1]
- 6 (a) any **three** from:  
barium more reactive / forms ions more readily  
barium reacts with (cold) water, nickel does not  
barium more vigorous with acids  
nickel compounds coloured, barium compounds white  
nickel has more than one oxidation state, barium has one  
nickel / nickel compounds catalysts, barium / barium compounds not catalysts  
nickel forms complex ions, barium does not [3]
- (b) (i) forward reaction favoured by low temperatures / reverse reaction favoured by high temperatures / heat [1]  
exothermic [1]
- (ii) products / RHS [1]  
has fewer moles / molecules / smaller volume / ORA [1]
- (iii) do not react or left behind / left at 60°C [1]

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- (iv) electrolysis [1]  
cathode (pure) nickel [1]  
anode impure nickel [1]  
electrolyte is a soluble nickel salt [1]
- 7 (a) correct method shown  
i.e.  $126/14 (= 9)$  or  $14x = 126$  or  $x = 9$  or  $(12 \times 9) + 18 = 126$  [1]  
 $C_9H_{18}$   
**note:** correct formula only = 1 [1]
- (b) (i) all hydrogen atoms 1bp [1]  
C—C bond atoms 1bp [1]  
C=C 2 bp [1]
- (ii) correct repeat unit [1]  
continuation [1]
- (iii) bonds broken  
H-H +436 (kJ/mol) C=C +610 = +1046 (kJ/mol) [1]  
bonds formed  
2C-H  $-415 \times 2$  kJ/mol C-C  $-346 = -1176$  (kJ/mol) [1]  
 $-130$  kJ/mol / more energy released than absorbed [1]  
**or:**  
bonds broken  
3882 (kJ/mol) [1]  
bonds formed  
4012 (kJ/mol) [1]  
 $-130$  kJ/mol / more energy released than absorbed [1]  
**allow:** ecf for final mark as long as the answer is not positive  
**note:** units not necessary
- (c) (i) butan-1-ol or butan-2-ol or butanol [1]
- (ii)  $CH_3-CH_2-CH(Br)-CH_2Br$  [2]  
 $C_4H_8Br_2 = 1$   
**note:** any other dibromobutane = 0
- (iii) HI [1]